

**CHAPTER 5 REVIEW***The Periodic Law***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1.   c   In the modern periodic table, elements are ordered
  - (a) according to decreasing atomic mass.
  - (b) according to Mendeleev's original design.
  - (c) according to increasing atomic number.
  - (d) based on when they were discovered.
2.   d   Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing
  - (a) density.
  - (b) reactivity.
  - (c) atomic number.
  - (d) atomic mass.
3.   b   The modern periodic law states that
  - (a) no two electrons with the same spin can be found in the same place in an atom.
  - (b) the physical and chemical properties of an element are functions of its atomic number.
  - (c) electrons exhibit properties of both particles and waves.
  - (d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
4.   c   The discovery of the noble gases changed Mendeleev's periodic table by adding a new
  - (a) period.
  - (b) series.
  - (c) group.
  - (d) level.
5.   d   The most distinctive property of the noble gases is that they are
  - (a) metallic.
  - (b) radioactive.
  - (c) metalloid.
  - (d) largely unreactive.
6.   c   Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
  - (a) 4.
  - (b) 10.
  - (c) 11.
  - (d) 18.
7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.
  - 9   a. How many protons are in this atom?
  - 10  b. How many neutrons are in this atom?
  - ${}^{19}_{9}\text{F}$  c. What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

# The Periodic Law Section Quiz

**Scott C. Dulebohn**



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