

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

# CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## States of Matter

### Section 13.1 Gases

In your textbook, read about the kinetic molecular theory.

Complete each statement.

1. The kinetic molecular theory describes the behavior of gases in terms of particles in motion.
2. The kinetic molecular theory makes the following assumptions:
  - a. In a sample of a gas, the volume of the gas particles themselves is very small compared to the volume of the sample.
  - b. Because gas particles are far apart, there are no significant attractive or repulsive forces between gas particles.
  - c. Gas particles are in constant and random motion.
  - d. The collisions between gas particles are elastic; that is, no kinetic energy is lost.
3. The kinetic energy of a particle is represented by the equation  $KE = \frac{1}{2}mv^2$ .
4. Temperature is a measure of the average kinetic energy of the particles in a sample of matter.

In your textbook, read about explaining the behavior of gases.

For each statement below, write true or false.

- |              |  |
|--------------|--|
| <u>True</u>  | 1. Gases are less dense than solids because there is a lot of space between the particles of a gas.                              |
| <u>True</u>  | 2. The random motion of gas particles causes a gas to expand until it fills its container.                                       |
| <u>False</u> | 3. The density of a gas decreases as it is compressed.   |
| <u>True</u>  | 4. A gas can flow into a space occupied by another gas.  |
| <u>True</u>  | 5. The diffusion of a gas is caused by the random motion of the particles of the gas.  |
| <u>False</u> | 6. Lighter gas particles diffuse less rapidly than do heavier gas particles.   |
| <u>True</u>  | 7. During effusion, a gas escapes through a tiny opening into a vacuum.  |
| <u>False</u> | 8. Graham's law of effusion states that the rate of effusion for a gas is directly related to the square root of its molar mass. |

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# CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY

## Section 13.2 continued

In your textbook, read about gas pressure.

Circle the letter of the choice that best completes the statement or answers the question.

9. Pressure is defined as force per unit c. area.
10. What is an instrument designed to measure atmospheric pressure? c. sphygmomanometer
11. The height of the liquid in a barometer is affected by all of the following EXCEPT the b. density of the liquid in the column.
12. The pressure of the gas in a manometer is directly related to which of the following quantities? d. height of the mercury column in the open end arm
13. One atmosphere is equal to a pressure of c. 101.3 kPa.
14. The partial pressure of a gas depends on all of the following EXCEPT the d. temperature of the gas.
15. The pressure of a sample of air in a manometer is 102.1 kPa. What is the partial pressure of nitrogen ( $N_2$ ) in the sample if the combined partial pressures of the other gases is 22.4 kPa? b. 79.7 kPa

Use the figure to answer the following questions.

16. What instrument is illustrated in the figure? barometer
17. Who invented this instrument? Evangelista Torricelli
18. What are the two opposing forces that control the height of the mercury in the column? gravity and atmospheric pressure
19. What does it mean when the level of mercury drops in the column? The atmospheric pressure has increased.



# Study Guide Matter And Change Answers

**Douglas Brandt**



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### **Table of Contents Study Guide Matter And Change Answers**

1. Understanding the eBook Study Guide Matter And Change Answers
  - The Rise of Digital Reading Study Guide Matter And Change Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Matter And Change Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide Matter And Change Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Matter And Change Answers
  - Personalized Recommendations

- Study Guide Matter And Change Answers User Reviews and Ratings
- Study Guide Matter And Change Answers and Bestseller Lists
- 5. Accessing Study Guide Matter And Change Answers Free and Paid eBooks
  - Study Guide Matter And Change Answers Public Domain eBooks
  - Study Guide Matter And Change Answers eBook Subscription Services
  - Study Guide Matter And Change Answers Budget-Friendly Options
- 6. Navigating Study Guide Matter And Change Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide Matter And Change Answers Compatibility with Devices
  - Study Guide Matter And Change Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide Matter And Change Answers
  - Highlighting and Note-Taking Study Guide Matter And Change Answers
  - Interactive Elements Study Guide Matter And Change Answers
- 8. Staying Engaged with Study Guide Matter And Change Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide Matter And Change Answers
- 9. Balancing eBooks and Physical Books Study Guide Matter And Change Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide Matter And Change Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Study Guide Matter And Change Answers
  - Setting Reading Goals Study Guide Matter And Change Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Study Guide Matter And Change Answers
  - Fact-Checking eBook Content of Study Guide Matter And Change Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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