

Phase Change Worksheet – Answer Sheet

- 1) A 12 oz. can of soda weighs about 450 grams. How many joules are released when a can of soda is cooled from 25 degrees Celsius (room temperature) to 4 degrees Celsius (the temperature of a refrigerator). The specific heat of liquid water is $4.184 \text{ J/g}^\circ\text{C}$. **39.5 kJ**
- 2) How many joules are required to heat 250 grams of liquid water from 0°C to 100°C ? **104.5 kJ**
- 3) How many joules are required to melt 100 grams of water? The heat of fusion of water is 334 J/g . **33.4 kJ**
- 4) How many joules are required to heat 100 grams of water? The heat of vaporization of water is 2260 J/g . **226 kJ**
- 5) How many joules are required to heat 100 grams of water from 25°C to 100°C ? The specific heat of water is $4.184 \text{ J/g}^\circ\text{C}$. **104.5 kJ**
- 6) How many joules are given off when 100 grams of water are cooled from 25°C to 10°C ? The specific heat of water is $4.184 \text{ J/g}^\circ\text{C}$. **104.5 kJ**
- 7) How many joules are required to heat 10 grams of water from 25°C to 100°C ? The specific heat of water is $4.184 \text{ J/g}^\circ\text{C}$. **33.6 kJ**
- 8) How many joules are required to heat a 100 gram cup of water 100 degrees from 0°C to the temperature of an over-heated refrigerator 100°C . The highest ambient temperature water is measured at 100°C . **418.4 kJ**
- 9) How many joules are released when 100 grams of liquid are cooled from 100°C to the ambient temperature of an over-heated refrigerator 100°C . The ambient temperature is 100°C . **0 kJ**
- 10) How many joules are required to heat the temperature of 100 grams of water from 100°C to the ambient temperature of water 100°C . The ambient temperature of water is 100°C . **0 kJ**

Study Guide Phase Changes Answer Key

Y Pai



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