

## REDOX REACTION

## Solved Examples

Example - 6 :



Sol. Valency factor of base = 1

Here, two molecule of NaOH replaced  $2\text{H}^+$  ion from the  $\text{H}_2\text{SO}_4$ . Therefore, each molecule of NaOH replaced only one  $\text{H}^+$  ion of acid, so v.f. = 1.

Q. v. f. for acid is the number of  $\text{OH}^-$  replaced from the base by each molecule of acid.

## Solved Examples

Example - 7 :



Sol. Valency factor of acid = 1

Here, one molecule of  $\text{H}_2\text{SO}_4$  replaced one  $\text{OH}^-$  from NaOH. Therefore, valency factor for  $\text{H}_2\text{SO}_4$  is one.

$$\therefore \text{Eq. wt. of } \text{H}_2\text{SO}_4 = \frac{\text{Mol. wt.}}{1}$$

Q. Salts :

(a) In non-reacting condition

Q. Valency factor = Total number of positive charge or negative charge present in the compound.

## Solved Examples

Example - 8 :

Sol.	$\text{Na}_2\text{CO}_3$	$\text{Fe}_2(\text{SO}_4)_3$	$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$
	v.f.	2	2
Eq. wt.	$\frac{M}{2}$	$\frac{M}{6}$	$\frac{M}{2}$

(b) In reacting condition

## Solved Examples

Example - 9 :



Sol. It is an acid base reaction, therefore valency factor for  $\text{Na}_2\text{CO}_3$  is one while in non-reacting conditions, it will be two.

(c) Equivalent weight of oxidising / reducing agents in a redox reaction

In case of redox change, v.f. = Total change in oxidation number per molecule.

# Study Guide Redox Reaction Answers

**Donald W. Shive**

## **Study Guide Redox Reaction Answers:**

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text. This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises. New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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**Study Guide and Full Solutions Manual** Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

Study Guide for Chang's Chemistry Kenneth W. Watkins, 1988

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