

METALS

A metal is an element which has the ability to lose electron(s) and become positively charged. There are about 105 different elements. Of these 84 are metals and 21 are non metals.

GENERAL PROPERTIES

Metals generally have the following properties.

- i. Have high melting and boiling points
- ii. Are good conductors of heat and electricity
- iii. Have a high density – metal atoms are packed very closely
- iv. Are solids at rtp, except for mercury which is a liquid.
- v. Are malleable i.e. can be beaten into sheets
- vi. Are ductile i.e. can be made into wires.
- vii. Are sonorous i.e. produce sound when beaten
- viii. Are lustrous i.e. can be polished or have a shining surface.
- ix. They are strong under tension and compression. That means they can withstand stretching and crushing without breaking.
- x. React with oxygen to form basic oxides
- xi. Form positively charged ions (Cations) by losing electrons.

WHY METALS ARE USED AS ALLOYS

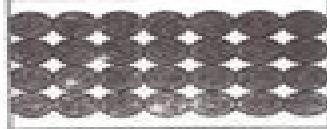
Pure metals are usually too soft and weak for most uses. A gold chain is never made of pure gold, as the chain would be too soft and would break easily. Instead, it is alloyed with other elements to make it more shiny and stronger.

Metals are mainly used as alloys for the following reasons

- To improve on their properties.
- To improve on their appearance.
- To make metals to be less brittle.
- To Make metals strong and harder.

A comparison between pure metals and alloys is shown below.

| Pure metal | Alloy |
|--|---|
| Are usually soft and can easily bend | Are strong and hard, not easily bent |
| Have higher melting points | Have lower melting points than the melting points of their components |
| Atoms can slide over each other easily as atoms are of the same size | Atoms cannot slide over each other easily as foreign atoms are of different size and disrupt the orderly arrangement of the metal atoms |


Diagrammatic representation of a pure metal


Diagrammatic representation of an alloy

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Julie R. Frentrup

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