

Molecular Polarity Table






A molecule is non polar if all surrounding atoms are same.

A molecule is polar if surrounding atoms are different.

A molecule is polar if central atom has one or more lone pair of electrons.

BP: Bond pair LP: Lone pair A: Central atom X: Surrounding atoms



Number of electron groups	Type of electron pairs	VSEPR notation	Name of Molecular shape	Example	Polarity
2	2 BP	AX_2	$X-A-X$ linear	BeF_2	Non-polar
3	3 BP	AX_3	 trigonal planar	BF_3	Non-polar
3	2 BP, 1 LP	AX_2E	 angular	$SnCl_2$	Polar
4	4 BP	AX_4	 tetrahedral	CF_4	Non-polar
4	3 BP, 1 LP	AX_3E	 trigonal pyramidal	PCl_3	Polar
4	2BP, 2LP	AX_2E_2	 angular	H_2S	Polar

Study Guide Molecular Shape Answer Key

Seager/Slabaugh



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