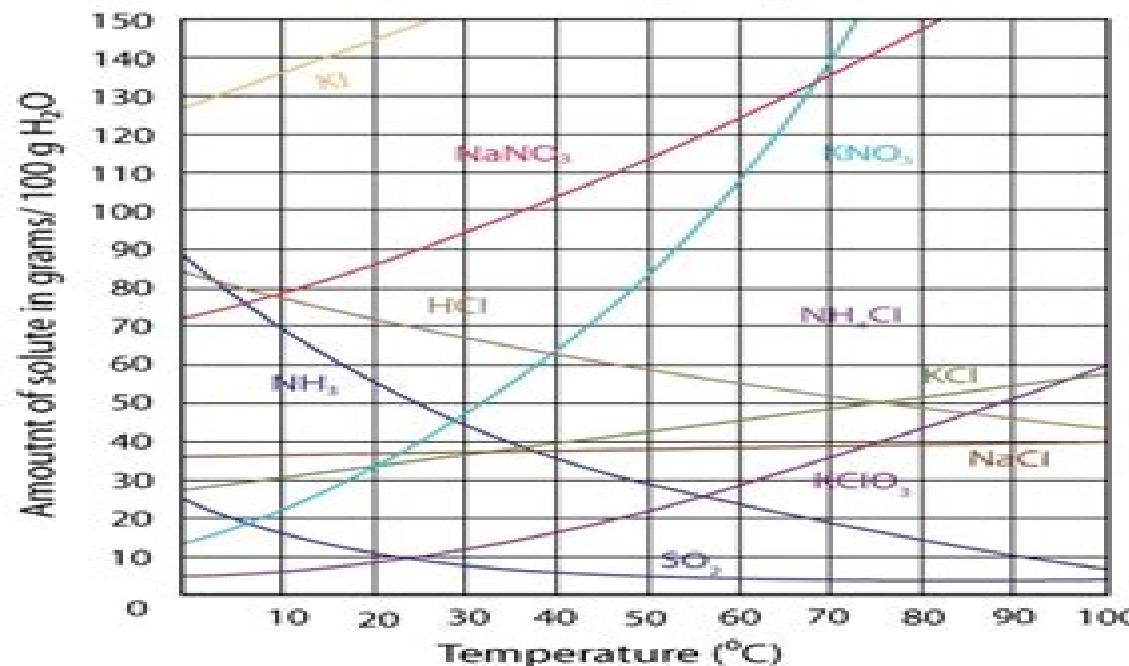


Interpreting Solubility Curve

1. Define solubility. _____

2. Using the solubility curve below, answer the given questions.



I. What is the solvent in this solubility curve? _____

II. Saturation points are given per how much of the solvent? _____

III. What is the saturation point of NaNO₃ at 10°C, 40°C, & 80°C?

10°C = _____ 40°C = _____ 80°C = _____

IV. How much NH₃ can you dissolve in water at 10°C, 30°C, & 90°C?

10°C = _____ 30°C = _____ 90°C = _____

V. Which salt is the least soluble in water? _____

VI. How many grams of potassium chloride can be dissolved in 200 g of water at 80 °C? _____

VII. At 40 °C, how much potassium nitrate will be dissolved in 300 g of water? _____

VIII. Which salt shows the least change in solubility from 0° to 100° C? _____

IX. Which substance shows a decrease in solubility from 0° to 100° C? _____

X. At 30 °C, 90 g of sodium nitrate is dissolved in 100 g water. Is this solution saturated, unsaturated, or supersaturated? _____

Study Guide Solubility Graph Key

**David Richerson,David W.
Richerson,William Edward Lee**

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