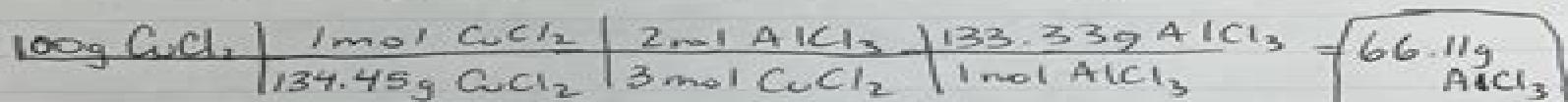
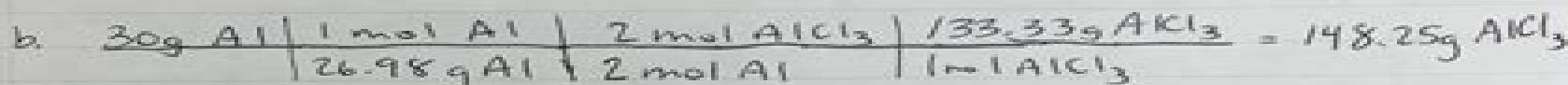
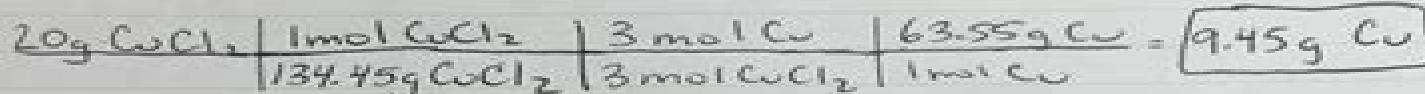
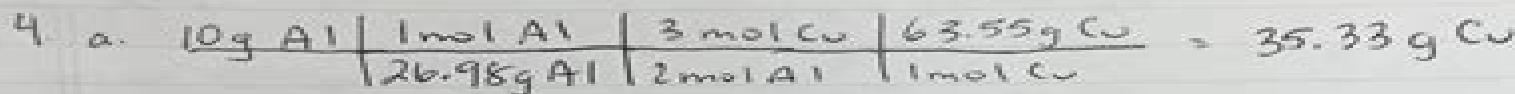


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~~1000000000000~~ $\frac{10 \text{ g}}{17.67 \text{ g}} \times 100 = 56.6\%$



~~1000000000000~~ $\frac{5 \text{ g}}{6.61 \text{ g}} \times 100 = 75.6\%$

5.

- TRUE
- FALSE
- $\text{AgNO}_3 \rightarrow 169.91 \text{ g/mol}$
- $6.02 \times 10^{23} \text{ particles} = 1 \text{ mole of particles}$

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Alan S. Wingrove, Robert L. Caret

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