

STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

STOICHIOMETRY

- Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.
- Stoichiometry is possible because of the Law of Conservation of Mass.
- The Law of Conservation of Mass states that the amount of matter (atoms) in a reaction remains equal to the amount of matter present after the reaction occurs.
- This law tells us that we can build ratios between reactants and products that allow us to calculate amounts of reactants.

Using Ratios

Start with the balanced chemical equation for the reaction. Use the coefficients to find the mole ratios between the reactants and products of the chemical reaction. Use the mole ratios to find the amount of excess reactant.



Percent Yield

Actual Yield	100% = Theoretical Yield
100 g	100 g = 100 g
200 g	200 g = 200 g
300 g	300 g = 300 g
400 g	400 g = 400 g
500 g	500 g = 500 g

Example: Finding the Limiting Reactant

No Excess Reactant

$\text{CaH}_2 + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

10.00 g of CaH_2 , react with 20.00 g of H_2O . Which is the limiting reactant and what is the actual mass of Ca(OH)_2 that can be formed?

Actual Yield	100% = Theoretical Yield
10.00 g	10.00 g = 10.00 g
20.00 g	20.00 g = 20.00 g
30.00 g	30.00 g = 30.00 g
40.00 g	40.00 g = 40.00 g
50.00 g	50.00 g = 50.00 g

Using Mole Ratios

Start with the balanced chemical equation. Use the coefficients to find the mole ratios between the reactants and products of the chemical reaction.

No Excess Reactant

How many moles of H_2 can be produced from 10.00 g of the reactant CaH_2 ?

How many moles of Ca(OH)_2 will form and H_2 is the limiting reactant?

$\text{CaH}_2 + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

A SET OF CHEMISTRY NOTES
MODULAR NOTES: USE THE PAGES
YOU NEED!

PRACTICE PROBLEMS AND
WORKED ANSWER KEYS INCLUDED

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Paul B. Kelter

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14

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