

# STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

## STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass.
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and product that allow us to calculate amounts of each.

## Using Ratios

Reactant and product amounts are related in a predictable way.

Each coefficient in the balanced equation can be used to define a ratio between the individual reactants and products of the chemical reaction. For the next example, let's use only moles of chemicals.

Consider the reaction:

$$2\text{LiCl} + \text{MgSO}_4 \rightarrow \text{Li}_2\text{SO}_4 + \text{MgCl}_2$$

We can write down ratios:

moles of LiCl	$\frac{2\text{LiCl}}{2\text{LiCl}}$	$\frac{2\text{LiCl}}{1\text{MgSO}_4}$
moles of MgSO <sub>4</sub>	$\frac{1\text{MgSO}_4}{2\text{LiCl}}$	$\frac{1\text{MgSO}_4}{1\text{MgSO}_4}$
moles of Li <sub>2</sub> SO <sub>4</sub>	$\frac{1\text{Li}_2\text{SO}_4}{2\text{LiCl}}$	$\frac{1\text{Li}_2\text{SO}_4}{1\text{MgSO}_4}$
moles of MgCl <sub>2</sub>	$\frac{1\text{MgCl}_2}{2\text{LiCl}}$	$\frac{1\text{MgCl}_2}{1\text{MgSO}_4}$
moles of LiCl	$\frac{2\text{LiCl}}{1\text{MgSO}_4}$	$\frac{2\text{LiCl}}{1\text{MgCl}_2}$
moles of MgSO <sub>4</sub>	$\frac{1\text{MgSO}_4}{2\text{LiCl}}$	$\frac{1\text{MgSO}_4}{1\text{MgCl}_2}$

## Example: Finding the Limiting Reactant in the reaction:

$2\text{LiCl} + \text{MgSO}_4 \rightarrow \text{Li}_2\text{SO}_4 + \text{MgCl}_2$  (aq)

5.37 grams of H<sub>2</sub>O, react with 5.43 grams of Mg. What is the limiting reactant and what mass of MgSO<sub>4</sub> can be formed?

First we need to find  $\frac{\text{moles of LiCl}}{\text{moles of MgSO}_4}$  and  $\frac{\text{moles of MgSO}_4}{\text{moles of LiCl}}$ .

For the reaction, then  $\frac{2\text{LiCl}}{1\text{MgSO}_4}$  should be  $\frac{\text{moles of LiCl}}{\text{moles of MgSO}_4}$ .

$\frac{5.37\text{g LiCl}}{42.39\text{g/mol LiCl}} \times \frac{1\text{MgSO}_4}{2\text{LiCl}} = \frac{0.063\text{MgSO}_4}{1}$  moles

$\frac{5.43\text{g Mg}}{24.31\text{g/mol Mg}} \times \frac{1\text{MgSO}_4}{1\text{Mg}} = \frac{0.223\text{MgSO}_4}{1}$  moles

0.064 grams of MgSO<sub>4</sub> will form and H<sub>2</sub>O is the limiting reactant.

## Using Mole Ratios

From the last page, we looked at the reaction:

$$2\text{Mg} + \text{MgSO}_4 \rightarrow 3\text{Mg} + \text{MgO}$$

We also defined ratios of moles of each substance using the coefficients of the balanced equation.

Now let's use the ratio we defined:

How many moles of Mg can be produced from 4 moles of LiCl in the reaction above?

We know the ratio of moles of Mg to moles of Li, whiching to be the given number of moles of Li. By writing this ratio into the denominator, we eliminate the units moles of Li and are left with moles of Mg.

4 moles  $\times \frac{1\text{Mg}}{2\text{LiCl}} = 2\text{ moles Mg}$

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# Study Guide Stoichiometry Limiting Reactant Percent Yield

**Paul B. Kelter**



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prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems      **Study Guide for Chang's Chemistry** Kenneth W. Watkins, 1988

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