

# SOLUBILITY AND SOLUBILITY PRODUCT



## Solubility and Solubility Product

| Solubility vs Solubility Product |  |
|----------------------------------|--|
| <b>DEFINITION</b>                | <b>Solubility</b><br>Property of a substance to dissolve in a solvent.   |
| <b>THEORY</b>                    | <b>Solubility product</b><br>Mathematical description of the equilibrium between dissolved and undissolved species of a substance in solution. |
| <b>FACTORS</b>                   | <b>Solubility product</b><br>Describes the dissolution of a substance in a solvent.  |

Solubility expresses the maximum concentration of solute in a solution. The solution with the maximum solute concentration is called a saturated solution. In a saturated solution, an equilibrium occurs so that it has an equilibrium constant called the solubility product constant ( $K_{sp}$ ).

- THE **SOLUBILITY PRODUCT ( $K_{sp}$ )** is the product of the concentrations of ions in a saturated solution raised to the power of the coefficient at a specific temperature. For example, in the electrolyte compound  $L_xZ_y(s)$ , the dissociation equation for the saturated solution is as follows.



$$K_{sp} = [L^{x+}]^a [Z^{x-}]^b$$

### **$K_{sp}$ RELATIONSHIP WITH SOLUBILITY (S)**

- If the  $K_{sp}$  value is large, then the solubility value (s) is also large  $\rightarrow$  the substance dissolves easily
- If the  $K_{sp}$  value is small, then the solubility value (s) is also small  $\rightarrow$  the substance is difficult to dissolve

$$K_{sp} = (n - 1)^{n - 1} \times S^n \quad \text{or} \quad K_{sp} = ((a+b) - 1)^{(a+b) - 1} \times S^{(a+b)}$$

# Study Guide Solubility And Concentration

**James E. Brady, John R. Holum**

## **Study Guide Solubility And Concentration:**

Fundamentals of Chemistry, Study Guide James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics. Emphasizes descriptive chemistry, chemical reactions and properties while maintaining a solid treatment of chemical principles. Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises. Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered. Chapters include frequent margin comments, figures and photographs.

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