

Stoichiometry and The Mole – Review Guide

Remember:

- 1 mole = 6.02×10^{23} atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



- 1) The process used to predict the amount of a product that will be formed is: stochiometric calculations
- 2) If you know the mass of a substance, you can find the number of moles by dividing it by the molar mass of substance.
- 3) If you know the number of particles (atoms) of a substance, you can find the number of moles by dividing it by Avogadro's Number.

Avogadro's Number

- 4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A \rightarrow Moles A \rightarrow Moles B \rightarrow Grams B

- 5) The essential stoichiometry problems involved in predicting from one chemical to another is:

Moles \rightarrow moles \rightarrow Moles

- 6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?

Chemical B

- 7) What is a mole? A unit of measurement used to count atoms/molecules.

- 8) What is Avogadro's Number? 6.02×10^{23}

- 9) What does Avogadro's number tell us? The number of atoms or molecules in a mole.

Study Guide The Mole With Answers

Robinson

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Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Study Guide with Selected Solutions Karen Timberlake,2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to

promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

Chemistry, Student Study Guide James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a

metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH4 from which nanotubes and buckyballs can be made

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