

12**STOICHIOMETRY****Chapter Test A****A. Matching**

Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line.

Column A

- _____ 1. actual yield
- _____ 2. limiting reagent
- _____ 3. theoretical yield
- _____ 4. stoichiometry
- _____ 5. percent yield
- _____ 6. excess reagent

Column B

- a. the ratio of the actual yield to the theoretical yield, expressed as a percentage
- b. the amount of product formed when a reaction is carried out in the laboratory
- c. the reactant that determines the amount of product that can be formed in a reaction
- d. the reactant that is not completely used up in a reaction
- e. the calculated amount of product that might be formed during a reaction
- f. the calculation of quantities in chemical equations

B. Multiple Choice

Choose the best answer and write its letter on the line.

- _____ 7. Which of these expressions is an incorrect interpretation of the balanced equation?
$$2S(s) + 3O_2(g) \rightarrow 2SO_3(g)$$
 - a. 2 atoms S + 3 molecules O₂ → 2 molecules SO₃
 - b. 2 g S + 3 g O₂ → 2 g SO₃
 - c. 2 mol S + 3 mol O₂ → 2 mol SO₃
 - d. none of the above
- _____ 8. In a chemical reaction, the mass of the products
 - a. is less than the mass of the reactants.
 - b. is greater than the mass of the reactants.
 - c. is equal to the mass of the reactants.
 - d. has no relationship to the mass of the reactants.
- _____ 9. How many liters of oxygen are required to react completely with 1.2 liters of hydrogen to form water?
$$2H_2(g) + O_2(g) \rightarrow 2H_2O(g)$$
 - a. 1.2 L
 - b. 0.6 L
 - c. 2.4 L
 - d. 4.8 L

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