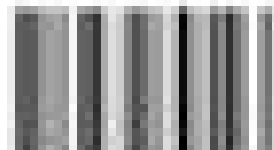


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subsequent  $\Delta h$  values  
 $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{O}_2 \text{ g}$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O g}$   $\Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6 \text{ g}$

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 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2\text{O l}$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O l}$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{H}_2\text{O l}$   $\Delta h = 285 \text{ kJ}$  mole

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_3\text{O}^+$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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