

PROBLEM SET 1.2. KINETICS AND DIFFUSION

ANSWER KEY

1.
 - A. The empirical formula of glucose is $C_6H_{12}O_6$. What is its molecular weight?
 $C_6H_{12}O_6$ should have a molecular weight of $(6 \times 12) + (12 \times 1) + (6 \times 16) = 72 + 12 + 96 =$
 180 g mol^{-1}
 - B. Isotonic glucose is 5% glucose w/v. How much glucose would we need to make 100 mL of isotonic glucose?
 5 g% means 5 g per dL; so we need **5 g** of glucose to make up this 100mL of isotonic glucose.
2.
 - A. You need to make 250 mL of a stock solution of 0.1 M Na_2ATP . Its formula weight is 605.2 g mol^{-1} . How much Na_2ATP should you weigh out?
 $\text{Amount of ATP} = C \times V = 0.1 \text{ mol L}^{-1} \times 0.25 \text{ L} = 0.025 \text{ mol}$
 $0.025 \text{ mol} \times 605.2 \text{ g mol}^{-1} = \textbf{15.13 g}$
 - B. Your advisor is skeptical of your abilities. He wants you to check out the 0.1 M ATP solution and tells you to do it spectrophotometrically. Spectrophotometry relies on the different abilities of chemicals to absorb light of specific wavelengths. A diagram of a spectrophotometer is shown below.

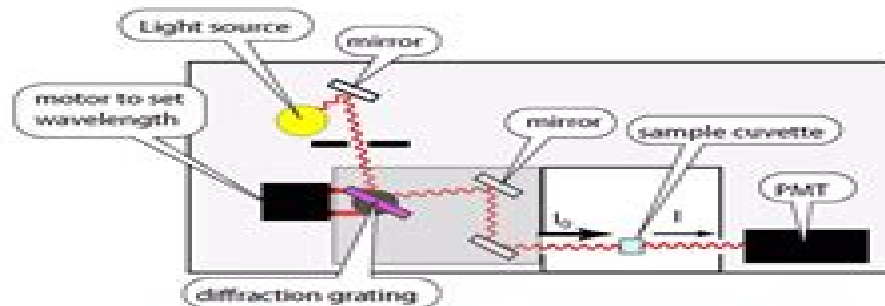


Fig. 1.PS2.1. Light path in a single-beam spectrophotometer. The view is from above. Light from a source is collimated (making a narrow beam) and passed through a monochromator that selects a narrow band of wavelength of light to be passed through the sample. A photomultiplier tube (PMT) detects the light and measures its intensity. Comparison of this intensity, I , to the intensity when the sample is missing, I_0 , allows calculation of the absorbance. Absorbance is recorded with time or as a function of wavelength.

At particular wavelengths, chemicals absorb light according to their chemical structure and their concentration. The law governing the absorption of light is the Beer-Lambert Law;

$$A = \epsilon C d$$

where A is the absorbance, ϵ is a constant that depends on the chemical and typically varies with the wavelength of light - it is the molar extinction coefficient and is in units of M^{-1} ; C is the concentration of the

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