

Solution Stoichiometry

1. What volume of a 0.35M AgNO_3 is required to completely react with 25 mL of a 0.24M NaCl solution?



Required: V of AgNO_3

Start with the other reactant

$$\frac{0.24 \text{ mol NaCl}}{1 \text{ L}} \cdot \frac{0.055 \text{ L NaCl}}{1 \text{ mol NaCl}} \cdot \frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}}$$

These cancel

$$\frac{0.35 \text{ mol AgNO}_3}{0.35 \text{ mol AgNO}_3} = 0.037 \text{ L} = 37.7 \text{ mL}$$

2. 37 mL of a 0.24M KI solution is required to completely react with 42 mL of a $\text{Pb}(\text{NO}_3)_2$ solution. What was the original concentration of the lead (II) Nitrate solution?



Given: 37 mL KI solution
 K_2 0.24M K_2 solution
 M_1 42 mL $\text{Pb}(\text{NO}_3)_2$

Required: original concentration of the lead

$$\frac{0.24 \text{ mol KI}}{1 \text{ L}} \cdot \frac{0.037 \text{ L}}{1} \cdot \frac{1 \text{ mol Pb}(\text{NO}_3)_2}{2 \text{ mol KI}} \cdot \frac{1}{0.042 \text{ L}} = 0.1277$$

$$2(42) \text{ M}_2 = 1(0.24)(37)$$

$$\text{M}_2 = 10.72 / 84$$

$$\text{M}_2 = 0.1277 \text{ M}$$

3. What mass of solid Magnesium Hydroxide can be produced if 15 mL of a 0.63M $\text{Mg}(\text{NO}_3)_2$ solution reacts completely with excess NaOH ?



Required: mass of solid Magnesium Hydroxide if 15 mL of a 0.63M $\text{Mg}(\text{NO}_3)_2$

Given: 15 mL $\text{Mg}(\text{NO}_3)_2$, 0.63M $\text{Mg}(\text{NO}_3)_2$

$$\frac{1 \text{ mol AgNO}_3}{1 \text{ mol NaCl}} \cdot \frac{0.63 \text{ mol Mg}(\text{NO}_3)_2}{1 \text{ L}} \cdot \frac{0.015 \text{ L}}{1} \cdot \frac{1 \text{ mol Mg(OH)}_2}{1 \text{ mol Mg}(\text{NO}_3)_2}$$

$$\cdot \frac{58.32 \text{ g}}{1 \text{ mol Mg(OH)}_2} = 1.65 \text{ g Mg(OH)}_2$$

4. What mass of solid Calcium Phosphate can be produced if 78mL of a 0.17M Sodium Phosphate solution is mixed with 38mL of a 0.21M Calcium Chloride solution?



Given:

$$\frac{78 \text{ mL Na}_3\text{PO}_4}{0.17 \text{ M Na}_3\text{PO}_4} \cdot \frac{38 \text{ mL CaCl}_2}{0.21 \text{ M CaCl}_2}$$

molar mass:

$$\text{Ca} = 2 \cdot 40.07 \quad \text{P} = 2 \cdot 30.97 \quad \text{O} = 2 \cdot 16 \quad \text{Na} = 22.99$$

Required:
 $\text{mass(g)} \text{ Ca}_3(\text{PO}_4)_2$

From: Na_3PO_4

$$\frac{0.17 \text{ mol Na}_3\text{PO}_4}{1 \text{ L}} \cdot \frac{0.078 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol Na}_3\text{PO}_4}$$

$$\cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} = 1.925 \text{ g Ca}_3(\text{PO}_4)_2$$

From CaCl_2

$$\frac{0.21 \text{ mol CaCl}_2}{1 \text{ L}} \cdot \frac{0.038 \text{ L}}{1} \cdot \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{2 \text{ mol CaCl}_2} \cdot \frac{310.18 \text{ g Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2}$$

$$= 1.4371 \text{ g Ca}_3(\text{PO}_4)_2$$

Solution Stoichiometry Practice Problems

Darrell D. Ebbing, Susan M. Schelble

Solution Stoichiometry Practice Problems:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshman in STEM fields.

Jacaranda Chemistry 1 VCE Units 1 and 2, learnON and Print Neale Taylor, Angela Stubbs, Robert Stokes, 2022-11-30 Developed by expert Victorian teachers for VCE students. The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities. Our expert author team of practising teachers and assessors ensures 100% coverage of the new VCE Chemistry Study Design 2023-2027.

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General Chemistry, Reactions First Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease. Using a friendly approach the text uses metaphors

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Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993 [Ebook: Chemistry: The Molecular Nature of Matter and Change](#) Silberberg,2015-01-16 [Ebook Chemistry The Molecular Nature of Matter and Change](#)

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